

Physical Science
Atomic History, Bohr Model, Electron Configuration,
And Periodic Table *Review*

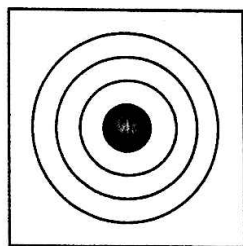
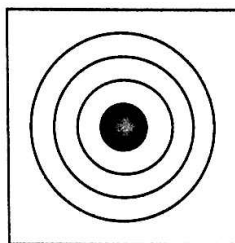
Atomic History

Study these people and what they did and/or discovered. I would write notes/ highlights next to their names.

1. Democritus
2. Dalton
3. Thompson
4. Bohr
5. Rutherford
6. Mendeleev

Bohr Model

Please draw the Bohr Model Diagram for the following **IN THE BOX** provided!

**Ar****N****Electron Configuration**

Please write out the Electron Configuration for the stable element listed below.

PUT ANSWERS IN THE CORRESPONDING BOXES. (Hint answers should be in numbers only)

8. Li_3^7

9. K_{19}^{40}

10. Mg_{12}^{24}

Periodic Table (Matching- Answers maybe used more than once.)

LOOK OVER PPTs and Notes

Label and Compute

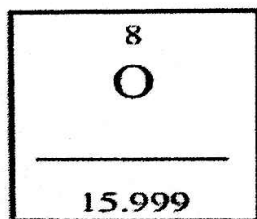
- _____ 11. # of Proton
- _____ 12. # of Neutrons
- _____ 13. # of Electrons
- _____ 14. Mass #
- _____ 15. Atomic #

16. → 3

Li

17. → 6.357

18. →



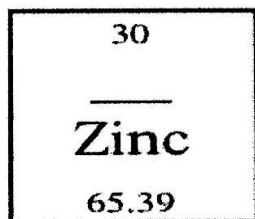
Atomic # = _____

Atomic Mass = _____

of Protons = _____

of Neutrons = _____

of Electrons = _____



Atomic # = _____

Atomic Mass = _____

of Protons = _____

of Neutrons = _____

of Electrons = _____

Atomic number equals
the number of

_____ or _____

Atomic mass equals
the number of

_____ + _____

Periodic Table																	
1 IA 1 H 1.00797	2 IIA											13 IIIA	14 IVA	15 VA	16 VIA	17 VIIA	18 VIIIA 2 He 4.0026
3 Li 6.939	4 Be 9.0122											5 B 10.811	6 C 12.0112	7 N 14.0067	8 O 15.9994	9 F 18.9984	10 Ne 20.179
11 Na 22.9898	12 Mg 24.305	3 IIIB	4 IVB	5 VB	6 VIB	7 VIIB	8	9 VIII	10	11 IB	12 IIB	13 Al 26.9815	14 Si 28.086	15 P 30.9738	16 S 32.064	17 Cl 35.453	18 Ar 39.948
19 K 39.102	20 Ca 40.08	21 Sc 44.956	22 Ti 47.90	23 V 50.942	24 Cr 51.996	25 Mn 54.9380	26 Fe 55.847	27 Co 58.9332	28 Ni 58.71	29 Cu 63.54	30 Zn 65.37	31 Ga 65.37	32 Ge 72.59	33 As 74.9216	34 Se 78.96	35 Br 79.909	36 Kr 83.80
37 Rb 85.47	38 Sr 87.62	39 Y 88.905	40 Zr 91.22	41 Nb 92.906	42 Mo 95.94	43 Tc [99]	44 Ru 101.07	45 Rh 102.905	46 Pd 106.4	47 Ag 107.870	48 Cd 112.40	49 In 114.82	50 Sn 118.69	51 Sb 121.75	52 Te 127.60	53 I 126.904	54 Xe 131.30
55 Cs 132.905	56 Ba 137.34	57 La 138.91	72 Hf 178.49	73 Ta 180.948	74 W 183.85	75 Re 186.2	76 Os 190.2	77 Ir 192.2	78 Pt 195.09	79 Au 196.967	80 Hg 200.59	81 Tl 204.37	82 Pb 207.19	83 Bi 208.980	84 Po [210]	85 At [210]	86 Rn [222]
87 Fr [223]	88 Ra [226]	89 Ac [227]	104 Ku [260]	105	106	107	108	109									

On the picture above.

1. Identify and label the Eight main families.
2. Which Families like to give up Electrons.
3. Which Families like to accept Electrons.
4. Circle the Staircase
5. Label/ identify Metals, Non-metals, and Metaliods.
6. How many Periods are on the Periodic table?
7. How many Columns are on the Periodic table?
8. What do all the elements in a Family have in common?
9. What do all the elements in a Period have in common?

LEWIS DOT DIAGRAMS:

THE DOTS AROUND AN ELEMENTS SYMBOL REPRESENT THE ELECTRONS IN THE OUTERMOST ORBITAL. THESE ELECTRONS ARE KNOWN AS _____.

DRAW THE LEWIS DOT DIAGRAM FOR THE FOLLOWING:

K Mg B P S F Ne